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## How To Make 12% Annually ... Forever

When pH=12,  $[OH^-] = 0.01 \text{ mol/L}$  ∴ in one litre ∴  $n = 0.01 \text{ mol}$  ∴  $M(\text{NaOH}) = 40 \text{ g/mol}$  So 0.4g ( $m = n \cdot M = 0.4 \text{ g}$ ) NaOH should be dissolved in one litre of solution. Hope these could help u~

## To make a solution of pH 12, what should the amount of ...

18) What volume (in mL) of 0.0887 M Mg I2 solution is needed to make 275.0 mL of 0.0224 M Mg I2 solution? A) 72.3 mL. B) 144 mL C) 69.4 mL D) 91.8 mL E) 10.9 mL

## Solved: 18) What Volume (in ML) Of 0.0887 M Mg I2 Solution ...

For example, to make 100 ml of 0.1 M CaCl<sub>2</sub> solution, use the previous formula to find out how much CaCl<sub>2</sub> you need: grams of CaCl<sub>2</sub> = (0.1) x (110.91) x (100) ÷ (1000) = 1.11 g; Now you can make your solution: dissolve 1.11 g of CaCl<sub>2</sub> in sufficient water to make 100 ml of solution. The amount of water needed will be slightly less than 100 ml.

## How to Make a Solution: Chemical, Molar and Weight Percent

Riley wants to make 100 mL of a 25% saline solution but only has access to 12% and 38% saline mixtures. Which of the following system of equations correctly describes this situation if X represents the amount of the 12% solution used, and y represents the amount of the 38% solution used? a.)  $0.12x + 0.38y = 0.25(100)$   
 $x + y = 100$  b.)  $0.38x + 0.12y = 100$   $x + y = 0.25(100)$  c.)  $0.38\% + 0.12y = 0.25(100)$

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Solved: Riley Wants To Make 100 ML Of A 25% Saline Solutio ...

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